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Science Division

The empirical formula for a compound is  $C_2H_5$  and its relative formula mass is 58.0. Deduce its molecular formula. Deduce its molecular formula. (Relative atomic masses:  $C = 12.0$ ,  $H = 1.0$ )

**Empirical formulae - Introducing chemical reactions - OCR ...**

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Empirical Formula Pogil Science Division Bellevue empirical formula. 2. Divide the molecular mass by the empirical mass. 3. Multiply all of the subscripts in the empirical formula by the resulting factor. Example Problem: What is the molecular formula for a material that is 92.2% C and 7.80% H by mass? The material has a molar mass of 52.0 g/mol. 1. Molecular mass /

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