

Chapter 9 Mixed Review Stoichiometry

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g)

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. Given the following equation: C₃H₄(g) + x O₂(g) → 3CO₂(g) + 2H₂O(g) a. What is the value of the coefficient x in this equation? b. What is the molar mass of C₃H₄? c. How many moles are in an 8.0 g sample of C₃H₄?

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CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: C₃H₄(g) + xO₂(g) → 3CO₂(g) + 2H₂O(g) a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C₃H₄? 2 mol O₂:1 mol H₂O

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Stoichiometry b. Theoretically, how many moles of NH₃ will be produced? PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g)

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Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry calculations start with a balanced chemical equation. This equation gives the ... 290 Chapter 9 DO NOT EDIT--Changes must be made through "File info" ...

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