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19. If 6.57 g of iron react with an excess of hydrochloric acid, HCl, then 11.2 g of iron(II) chloride are obtained in addition to hydrogen gas. Find the theoretical and percent yields.  $\text{Fe} + 2 \text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2$  20. If 5.45 g of potassium chlorate are decomposed to form potassium chloride, 1.75 g of oxygen gas are also given off.

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II—I Review and Reinforcement (continued) Solve each of the following problems as directed. Show all your work. 13. How many moles Of magnesium are required to react with 2.0 mol of hydrochloric acid? The equation for this reaction is  $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$  14. Aluminum reacts with HCl to produce aluminum chloride ( $\text{AlCl}_3$ ) and hydrogen ...

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11—2 Review and Reinforcement Solving Stoichiometry Problems If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line provided. 1. 2. 5. The major types of stoichiometry problems are mass-mass, mass-volume, and volume-volume.

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